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on the solubility of NaCl.
8. Explain how you might make a solution containing 42 g KCl dissolved in 100 g H₂O at a temperature of 40°C. What term describes this type of solution?
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1. Which solid has the lowest solubility at 60 °C?
2. Which solid has the same solubility as potassium nitrate at 52 °C?
3. Which solid has a solubility that changes

least with temperature?
4. Which solid would give a deposit of 20 g if a saturated solution in 100 g of water at 60 °C was cooled to 20 °C?
5.

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Table 14.1: Solubility of table salt as a function of water. In this graph 6 measurements of the solubility were made at different temperatures. Here, the solubility is the maximum amount of salt that can be dissolved in 100 grams of water, and any extra salt beyond that falls to the bottom of the container as a precipitate (solid).

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