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1. Aqueous solutions of acids have a sour taste. 2. Acids change the color of acid-base indicators. 3. Some acids react w/active metals to release hydrogen gas, H<sub>2</sub>. 4. Acids react with bases to produce salts and water. (salt- the reaction products are water and an ionic compound called a salt). 5. Some acids conduct electric current. 6 ...

What contains acetic acid? Vinegar What contains lactic acid?

sour milk What does phosphoric acid do to many carbonated beverages? It gives a tart flavor What foods contain citric acid? Lemons, oranges, grapefruits What contains malic acid? Apples What contains tartaric acid? Grapejuice What is the common name for Sodium Hydroxide NaOH? lye Acids were [...]

CH 1 Reading Assignment Modern Chemistry. CH 1 Vocabulary-New. CH 1 Mixed Questions. CH 1 Matter & Energy Vocabulary-New. CH 1 Classifying Matter-New. CH 1 Classification of Matter ... Chapter Fourteen Acids & Bases. Objectives: Chapter 14. Notes: Chapter 14 . Reference: CH 14 Acid Base Organizer.

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#### **CHAPTER 14 REVIEW Acids and Bases**

1. Aqueous solutions of acids have a sour taste. 2. Acids change the color of acid-base indicators and cause pH paper to become red. 3. Some acids react with active metals and release hydrogen gas. 4. Acids react with bases to produce salts and water. 5. Acids conduct electric current.

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3 is the Lewis acid because it accepts an electron pair in forming a covalent bond. It is not a Brønsted-Lowry acid, because it is not donating a proton. e. Which of the reactants is the Lewis base? Explain your answer. NF 3 is the Lewis base because it donates an electron pair in forming a covalent bond. 120 ACIDS AND BASES MODERN CHEMISTRY

## 14 Acids and Bases

What contains acetic acid? Vinegar What contains lactic acid? sour milk What does phosphoric acid do to many carbonated beverages? It gives a tart flavor What foods contain citric acid? Lemons, oranges, grapefruits What contains malic acid? Apples What contains tartaric acid? Grapejuice What is the common name for Sodium Hydroxide NaOH? lye Acids were [...]

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## and ...

Modern Chemistry 142 Chapter Test Name Class Date Chapter Test B, continued 28. Write the general equilibrium expression for the dissociation of an acid-base indicator that is a weak acid, HIn, and explain how this equilibrium determines the color of the indicator at a given pH. PART IV

## Assessment Chapter Test B - Baumapedia

These lecture presentations were designed for my high school Chemistry I Honors class. Students of high school and college general chemistry may find them useful as a supplement to their own class notes or as a review. Teachers, please feel free to use and modify them for your own classes.

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Modern Chemistry 121 Acids and Bases CHAPTER 14 REVIEW Acids and Bases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Answer the following questions according to the Brønsted-Lowry definitions of acids and bases: \_\_\_\_ a. What is the conjugate base of H<sub>2</sub>SO<sub>3</sub>? \_\_\_\_ b.

### CHAPTER 14 REVIEW Acids and Bases

Modern Chemistry Chapter 15 Test Review 25 multiple choice definition of self-ionization  $[H_3O^+][OH^-] = 1 \times 10^{-14}$  definition and formula for calculating  $pH = -\log[H_3O^+]$   $pH > 7$  is a base;  $pH < 7$  is an acid;  $pH = 7$  is neutral  $pH$  range is normally 0 to 14 calculate  $pH$  from  $[H_3O^+]$  calculate  $[H_3O^+]$  from  $pH$  definitions of indicators, transition ...

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Modern Chemistry Chapter 15 Test Review 25 multiple choice

definition of self-ionization  $[\text{H}_3\text{O}^+][\text{OH}^-] = 1 \times 10^{-14}$  definition and formula for calculating  $\text{pH} = -\log[\text{H}_3\text{O}^+]$   $\text{pH} > 7$  is a base;  $\text{pH} < 7$  is an acid;  $\text{pH} = 7$  is neutral  $\text{pH}$  range is normally 0 to 14 calculate  $\text{pH}$  from  $[\text{H}_3\text{O}^+]$  calculate  $[\text{H}_3\text{O}^+]$  from  $\text{pH}$  definitions of indicators, transition ...

Modern Chemistry 142 Chapter Test Name Class Date Chapter Test B, continued 28. Write the general equilibrium expression for the dissociation of an acid-base indicator that is a weak acid,  $\text{HIn}$ , and explain how this equilibrium determines the color of the indicator at a given  $\text{pH}$ . PART IV

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$\text{SO}_3^{2-}$  is the Lewis acid because it accepts an electron pair in forming a covalent bond. It is not a Brønsted-Lowry acid, because it is not donating a proton. e. Which of the reactants is the Lewis base? Explain your answer.  $\text{NF}_3$  is the Lewis base because it donates an electron pair in forming a covalent bond. 120 ACIDS AND BASES MODERN CHEMISTRY

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