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Equilibrium constants In order to have more information on the ion exchange process on the material surface, the constants of the ion exchange equilibria were determined For the ideal ion exchange model, the equilibrium reaction can be represented by the

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Name Formula Ka1 Ka2 Ka3 Acetic acid CH₃COOH (or HC₂H₃O₂) 1.8 * 10⁻⁵ Arsenic acid H₃AsO₄ 5.6 * 10⁻³ 1.0 * 10⁻⁷ 3.0 *

10-12 Arsenous acid H_3AsO_3 5.1×10^{-10}
Ascorbic

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For the ideal ion exchange model, the equilibrium reaction can be represented by the following equation: $M^{++} + H_2D$

$M^{++} + H_2D + K(2)$ where the bars mean the ions in the solid phase, and K is the equilibrium constant. It must be pointed out that this model failed to approximate the experimental data within the limits of their errors.

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DETERMINATION OF ION EXCHANGE EQUILIBRIUM CONSTANTS FOR ...

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Stability constants of complexes - Wikipedia

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ION EXCHANGE EQUILIBRIUM STUDY

USING STRONGLY BASIC ANION ...

The equilibrium constant for the formation of the complex ion from the hydrated ion is called the formation constant (Kf). The equilibrium constant expression for Kf has the same general form as any other equilibrium constant expression. In this case, the expression is as follows: $K_f = \frac{[\text{Cu}(\text{NH}_3)_4]^{2+}}{[\text{Cu}^{2+}][\text{NH}_3]^4} = 2.1 \times 10^{13}$

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The ion exchange isotherms at 302 K for $\text{Na}^+/\text{Cu}^{2+}$ and $\text{Cu}^{2+}/\text{Na}^+$ on zeolite A were determined for six total equivalent concentrations of the external solution, in the range 0.05–2.1 eq/L. Interpolated points from the curves fitted with different isotherms were used in the calculation of the selectivity coefficients. The activity coefficients in the external solution were calculated by means of ...

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ION EXCHANGE EQUILIBRIA IN BINARY AND TERNARY SYSTEMS

In contrast to an ordinary cation-exchange resin, the ion exchange behavior of Mg^{2+} and Ca^{2+} on the amphoteric ion-exchange resin showed a marked dependence on the kinds of salts: the distribution coefficients for the NaCl system were independent of the salt

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