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Equilibrium is reached again with the precipitation of more NaCl to compensate for the added Cl- ions by shifting to the left Over-concentrated KSCN + Fe2(NO3)3 solution Immersed in boiling water, the solution slowly turned blue Fe+3(aq) + SCN-(aq) FeSCN+2(aq) Removing heat

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Answer to Please answer the post lab questions in accordance to the data provided and the procedure at end Procedure Skip Navigation. Chegg home. Books. ... Name: Class/Lab Period: Exploring Equilibrium Data Table Part A. Complex-Ion Equilibrium of Iron(III) and Thiocyanate Ions Color of KSCN Color of Fe(NO)3 Solution Solution oranSe ...

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Exploring Equilibrium Lab Pre-Lab Questions 1. False, The reactions continue, but there is an equal balance of opposing reaction rates 2. False, the concentration of reactants and products remain constant with time. They don't necessarily have to be equal. 3. (a) When the humidity is low, the paper will be blue.

Color lab - Exploring Equilibrium Lab Pre-Lab Questions 1 ...

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Equilibrium POST-LAB QUESTIONS] 1. Write the chemical equation for the complex-ion equilibrium that results when excess chloride ion is added to an aqueous solution of cobalt chloride. Note the observed color of each complex ion underneath its chemical formula.

Chemistry Post-Lab Questions | Chemical Equilibrium | Chloride

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Chemistry at TPA: Exploring Equilibrium - It Works Both Ways

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Solved: Please Answer The Post Lab Questions In Accordance ...

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Equilibrium Lab | Chemical Equilibrium | Reaction Rate

I used "Reactions and Rates" in my regular chem class in this brief activity, which also includes a non-PhET, hands-on activity. My use of the sim was for conceptual purposes only; I was hoping to attack some misconceptions about equilibrium, activation energy, and temperature's affect on equilibrium position.

Exploring Equilibrium - PhET Contribution

This short activity is done after Le Chatelier's Principle is covered in class. Students have already done the first lab, "Exploring Equilibrium," where they investigated temperature changes a little bit; here, they test their ideas and predictions about temperature changes in light of Le Chatelier's Principle. Subject Chemistry

Exploring Equilibrium 2 (Le Chatelier's Principle) - PhET ...

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b) Describe the nature of dynamic equilibrium when small numbers of particles (such as 50, as compared to 6.022×10^{23}) are present. 4) Predict what will happen when the temperature is raised so it is NOT above the activation energy max but IS above the energy level of the products. 5) Raise the temperature as described.

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Exploring Equilibrium Mini Lab Part A

Chemistry Equilibrium Lab Question? Hi! Okay so I have a question about the calculations for an equilibrium lab I completed. I have the equation . $\text{Fe}^{3+} + \text{SCN}^{-1}$ makes FeSCN^{2+} I know that Fe^{3+} equals 2.25×10^{-5} M and SCN^{-1} equals 0.50 M. Would I simply add the 2 together to get FeSCN^{2+} ? That seems too easy but...

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Exploring Equilibrium Mini Lab Part B

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