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The relative masses were obtained by multiplying the atomic ratios and atomic mass-

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mol The following equation represents a laboratory preparation for oxygen gas:  $2KClO_3(s) \rightarrow 2KCl(s) + 3O_2(g)$  How many moles of  $O_2$  form if 3.0 mol of  $KClO_3$  are totally consumed?

Stoichiometry The study of quantitative relationships between the amounts of reactants used and products formed by a chemical reaction; based on the law of conservation of mass Actual Yield

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Stoichiometry is based on the law of conservation of mass. \_\_\_\_ 3. In any chemical reaction, the mass of the products is less than the mass of the ... Chemistry: Matter and Change 1 Study Guide. 1311 ... TEACHER GUIDE AND ANSWERS Study Guide - Chapter 11 - Stoichiometry Section 11.1 What is stoichiometry? 1. true 2. true 3. false 4. true

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Stoichiometry is the tool for answering these questions. Stoichiometry The study of quantitative relationships between the amounts of reactants used and amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass. Recall from Chapter 3 that the law states

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