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Chemistry Chapter 10 The Mole Section 10.4 Empirical and Molecular Formulas 1. The percent composition of a compound is the percent by mass of each

of the elements in a compound. 2.

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What is the mole? Chemists use the mole to count atoms, molecules, ions, and formula units. It is important to note that a mole always contains the same number of particles. For example, you know that a dozen eggs will always contain 12 eggs. Likewise, a mole will always contain 6.022×10^{23} particles. Although a mole of any substance is the same number, different substances have different molar masses. Start studying Chemistry: Chapter 10- THE MOLE. Learn vocabulary, terms, and more with flashcards, games, and other study tools.

The Mole | Introductory Chemistry mole of water is the water molecule, the

representative particle in a mole of copper is the copper atom, and the representative particle in a mole of sodium chloride is the formula unit. 310 Chapter 11 The Mole Figure 11-2 The amount of each substance shown is 6.02×10^{23} or one mole of representative particles. The representative particle for

The MoleThe Mole - Weebly Chapter 10 • The Mole 319 SStart-Up Activitiestart-Up Activities LAUNNCH CH LLabab How much is a mole? Counting large numbers of items is easier when you use counting units such as decades or dozens. Chemists use a counting unit called the mole. Procedure 1. Read and complete the lab safety form. 2. Select an item to measure, such as a paper clip, gum Over the next 3 week break, we are going to transition into studying the concept of the mole and how its used to measure quantities of chemical substances. You should create a new divider in your binder titled "Chapter 10: The Mole".

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A mole (mol) is a number of things equal to the number of atoms in exactly 12 g of carbon-12. Experimental measurements have determined that this number is very large: $1 \text{ mol} = 6.02214179 \times 10^{23}$ things Understand that a mole means a number of things, just like a dozen means a certain number of things—twelve, in the case of a dozen.

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Chapter 10

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