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Chapter 4 Atomic Structure

Dalton's Atomic Theory (experiment based!) 3) Atoms of different elements combine in simple whole-number ratios to form chemical compounds 4) In chemical reactions, atoms are combined, separated, or rearranged - but never changed into atoms of another element.

104 Chapter 4 • The Structure of the Atom Figure 4.2 In his book A New System of Chemical Philosophy, John Dalton presented his symbols for the elements known at that time and their possible combinations.

Chapters 4, 5 & 6. This unit's focus is on atomic structure. Students will cover concepts related to atomic mass, isotopes, electron configurations and periodicity.

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Dalton's Atomic Theory (experiment based!) 3) Atoms of different elements

combine in simple whole-number ratios to form chemical compounds 4) In chemical reactions, atoms are combined, separated, or rearranged - but never changed into atoms of another element. 1) All elements are composed of tiny indivisible particles called atoms

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atomic number: equals the number of protons in an atom of that element: mass number: the sum of the protons and neutrons in the nucleus of that atom: isotopes: atoms of the same element that have different numbers of neutrons and different mass numbers: energy levels: the possible energies that electrons in an atom can have: electron cloud

FPS chapter 4 : Sankovich - Nov - 201 2 4 test. Sunday Monday Tuesday Wednesday HW: Study for test Thursday Friday History Saturday Sept 17 Intro WS : Advisory Schedule. Welcome to SHHS. Mystery Box Safety Contract (pp7 - 8) (pp5 - 6) HW: Intro

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a vertical column of elements in the periodic table; the constituent elements of a group have similar chemical and physical properties.

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Chapter 4: The Structure of the Atom

Chapters 4, 5 & 6. This unit's focus is on atomic structure. Students will cover concepts related to atomic mass, isotopes, electron configurations and periodicity.

Unit 2 - Chapters 4, 5 & 6 - Mrs. Gignras' Chemistry Page

The atomic mass of an isotope is the sum of the number of protons and neutrons in its nucleus. The max. number of electrons in the first energy level of an atom is:2. in the second energy level is:8. in the third energy level is:18. A certain atom has 26 protons, 26 electrons, and 30 neutrons.

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