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atomic number: equals the number of protons in an atom of that element: mass number: the sum of the protons and neutrons in the nucleus of that atom: isotopes: atoms of the same element that have different numbers of neutrons and different mass numbers: energy levels: the possible energies that electrons in an atom can have: electron cloud

The atomic mass of an isotope is the sum of the number of protons and neutrons in its nucleus. The max. number of electrons in the first energy level of an atom is:2. in the second energy level is:8. in the third energy level is:18. A certain atom has 26 protons, 26 electrons, and 30 neutrons.

a vertical column of elements in the periodic table; the constituent elements of a group have similar chemical and physical properties.

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Dalton's Atomic Theory (experiment based!) 3) Atoms of different elements combine in simple whole-number ratios to form chemical compounds 4) In chemical reactions, atoms are combined, separated, or rearranged – but never changed into atoms of another element.

Chapters 4, 5 & 6. This unit's focus is on atomic structure. Students will cover concepts related to atomic mass, isotopes, electron configurations and periodicity.

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a vertical column of elements in the periodic table; the constituent elements of a group have similar chemical and physical properties.

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Chapter 4: Atomic Structure Worksheet . Answer the following questions, circle the best answer. 1) Which of the following are isotopes of each other? a) ^{14}C and ^{14}N b) ^3H and ^4He c) ^2H and ^1H d) none of these . 2) The net charge on an atom that has 13 protons, 12 neutrons, and 10 electrons is . a) $5+$ b) -3 c) $3+$ d) -9

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Dalton's Atomic Theory (experiment based!) 3) Atoms of different elements combine in simple whole-number ratios to form chemical compounds 4) In chemical reactions, atoms are combined, separated, or rearranged - but never changed into atoms of another element. 1) All elements are composed of tiny indivisible particles called atoms

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Chapter 4: The Structure of the Atom

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Chapter 4: The Structure of the Atom

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