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SECTION 2. SHORT ANSWER Answer the following questions in the space provided. 1. In cathode-ray tubes, the cathode ray is emitted from the negative electrode, which is called the \_\_\_\_\_ . 2.

Chapter 4 Arrangement Of Electrons In Atoms Section 3

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3 Atoms: The Building Blocks of Matter

Chapter Three [The Building Blocks of Matter]  
CHAPTER 4 REVIEW. Arrangement of Electrons in Atoms. SECTION 4.3. SHORT ANSWER Answer the following questions in the space provided. 1. State the Pauli exclusion principle, and use it to explain why electrons in the same orbital must have opposite spin states. 2. Explain the conditions under which the following orbital notation for helium is ...

CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

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Chapter 3 Review Atoms Section

1. Elements are made of particles that are very small known as atoms. 2. All of the atoms of the same element are duplicates of each other/identical. 3. The atoms of one element are different from those of any other type of element. 4. Atoms of one element can combine with atoms of different elements to make compounds.

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3 Atoms: The Building Blocks of Matter

1. All matter is composed of atoms. 2. Atoms of a given element contain all the same properties, and vice versa. 3. Atoms cannot be created, destroyed, or subdivided. 4. Atoms of different elements combine in simple whole numbered ratios, to form chemical compounds. 5. In chemical reactions, atoms are combined, separated. or rearranged.

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Chemistry- Chapter 3, Review Questions: Section 2 ...

3. State two principles from Dalton's atomic theory that have been revised as new information has become available. 4. The formation of water according to the equation.  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ . shows that 2 molecules (made of 4 atoms) of hydrogen and 1 molecule (made of 2 atoms) of oxygen produce 2 molecules of water.

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Chapter Four [Arrangement of Electrons in Atoms] Chapter Five [The Periodic Law] Chapter Six [Chemical Bonding] ... Section 2: Chapter review 3 and 6. Section 3: Chapter review 7 thru 16. Work practice problems 17, 18, 19, 21,22 and 23 (pp 89-90). Homework Answers . Problem Solving Packet.

Chapter Three [The Building Blocks of Matter]

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Section Quiz, continued Class Date ectra of atoms with how many 6. Bohr's model corr ctly explains the electron(s)? a one . two c. three d. four or more 7. A form of energy that exhibits wave behavior as it travels through space is a. microwave radiation. b. ultraviolet radiation. c. infrared radiation. I of the above rpDg n 8.

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CHAPTER 5 Electrons in Atoms Resource Manager Objectives Section Section 5.1 What You'll Learn Using the Photo Electrons in Atoms Have students review the following concepts Analysis Answers will vary.

CHAPTER 3 REVIEW . Atoms: The Building Blocks ofMatter. SHORT ANSWER . Answer the following questions in the space provided. 1. In cathode-ray tubes, the cathode ray is emitted from the negative electrode, which is called the . cathode 2. The smallest unit of an element that can exist either alone or in molecules containing.the

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