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Chapter 16 Thermal Energy and Heat Section 16.2 Heat and Thermodynamics (pages 479-483) This section discusses three kinds of thermal energy transfer and introduces the first, second, and third laws of thermodynamics. Reading Strategy (page 479) Build Vocabulary As you read this section, add definitions and examples to complete the table.

16.1 Thermal Energy and Matter Heat flows spontaneously from hot objects to cold objects. • Heat is the transfer of thermal energy from one object to another because of a temperature difference.

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Work and Heat Temperature is the measure of how hot or cold something is compared to a reference point. The Celsius scale has reference points of freezing and boiling points of water On the Kelvin scale the reference point is absolute zero Absolute Zero is the temperature at which molecules essentially stop (no kinetic energy)

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the study of conversions between thermal energy and other forms of energy: heat engine: any device that converts thermal energy into work: waste heat: thermal energy discharged into an area at a lower temperature without being converted into useful work: external combustion engine: a heat engine in which the fuel burns outside the engine

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Chapter 16 Thermal Energy and Heat Section 16.1 Thermal Energy and Matter (pages 474-478) This section defines heat and de-

scribes how work, temperature, and thermal energy are related to heat. Thermal expansion and contraction of materials is discussed, and uses of a calorimeter are explained. Reading Strategy (page 474)

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The quantity of energy transferred by heat from a body depends on a number of physical properties of the body and its surroundings. For a given substance, the rate at which thermal energy is transferred by conduction depends on temperature difference, cross-sectional area, and a thermal conductivity constant that is unique to the substance.

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