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Chapter 15 — Water & Aqueous System. the chaotic movement of colloidal particles, caused by collision with particles of the solvent in which they are dispersed. a compound that conducts an electric current when it is in an aqueous solution or in the molten state; all ionic compounds are electrolytes, but most covalent compounds are not.

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**SECTION 15.1 WATER AND ITS PROPERTIES (pages 445-449)**

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 Chapter 15 Water and Aqueous Systems. Chapter 15 "Water and Aqueous Systems". The Water Molecule: a Review.  
 • Water is a simple tri-atomic molecule, H<sub>2</sub>O.  
 • Each O-H bond is highly polar, because of the high electronegativity of

the oxygen (N, O, F, and Cl have high values) • bond angle of water = 105o.

### Chapter 15 Water and Aqueous Systems

Chapter 15 - Water and Aqueous Systems - 15.2 Homogeneous Aqueous Systems - 15.2 Lesson Check - Page 501: 11 Answer All ionic compounds are electrolytes because they dissociate into ions and conduct an electrical current in aqueous solution or in the molten state.

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