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Electrons in Atoms & Periodic Relationships Chapter 13-14 ...
12. The maximum # of electrons allowed in an energy level = 2n2 (where n is the level number) Ex: level 1 2(1) 2 = 2 electrons max. level 2 2(2) 2 = 8 electrons max. 13. Sublevels: s,p,d,f some people don’t forget. s sublevel and orbitals. s sublevels have 1 orbital. Each orbital can hold 2 electrons. Spherical shaped orbital
Section 13.1 Chapter 13 Electrons in Atoms z
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Chapter 13: Electrons in Atoms. The Bohr model limits electrons to specific circular

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Chapter 13 Electrons in Atoms. Adapted from notes by Stephen L. Cotton ©2006. Section 13.1 Models of the Atom. zOBJECTIVES: Summarize the development of atomic theory. Explain the significance of quantized energies of electrons as they relate to the quantum mechanical model of the atom.

Chapter 13 – Electrons in Atoms Chapter 13: 1 – 20, 23 – 25, 27, 31, 32, 34 – 38, 41, 45, 47, 48, 52 Section 13.1 – Models of the Atom Section Review 13.1 1. List in chronological order, a major contribution of each of these scientists to the understanding of the atom: proposed that all elements are composed of atoms.Dalton –

Chapter 13 Homework

Chapter 10 - States of Matter; Chapter 11 - Thermochemistry; Chapter 12 - The Behavior of Gases; Chapter 13 - Electrons in Atoms; Chapter 14 - Chemical Periodicity; Chapter 15 - Ionic Bonding and Ionic Compounds; Chapter 16 - Covalent Bonding; Chapter 17 - Water and Aqueous Systems; Chapter 18 - Solutions; Chapter 19 - Reaction Rates & Equilibrium
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Chapter 13 Electrons in Atoms

Chapter 13 Electrons in Atoms. when electrons occupy orbitals of equal energy, one electron enters each orbital until all orbitals contain one electron and all electrons will have parallel spins, must fill all boxes with one arrow first.

CHAPTER 13 Electrons in Atoms

Chapter 13: Electrons in Atoms

Chapter 13- Electrons in Atoms ★ Atomic model evolution: John Dalton- atom is solid indivisible mass, explains the nature of chemical reactions, discovery of subatomic particles JJ Thomson- discovered the electron → plum pudding model → negatively charged electrons stuck to positive mass → no info on protons and electrons or their location Ernest Rutherford- dense nucleus, rest empty ...

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Chapter 13: Electrons in Atoms The Evolution of Atomic Models Section 13.1 Describe an atom. What are the three main subatomic particles? Identify the relative electrical charges associated with each particle.

Quantum Mechanical Model The energy levels are not equally spaced like a ladder – they get closer the farther from the nucleus you go The higher the energy of the e-, the easier it leaves the atom

Bohr proposed a planetary model where electrons orbit the nucleus in an elliptical path much as planets orbit the sun-- earth orbits the sun so fast that it does not crash into the sun. Can only orbit at certain distances.

Chapter 13 & 14 Assignment & Problem Set 7. An atom of an element has two electrons in the first energy level and five electrons in the second energy level. Write the electron configuration for this atom and name the element.

Chapter 5 Electrons in Atoms . Name Date 11. The number of sublevels in an energy level is equal to the square of the principal quantum number of that energy level. ... 13. As many as four electrons can occupy the same orbital. 14. The Pauli exclusion principle states that an atomic orbital may describe at most two electrons.

chapters 13&14 - Chapter 13 Electrons in Atoms ...

Electrons move in circular orbits . around . the nucleus . at . fixed. energy. levels. Electrons are never between energy levels or energy shells. An electron must have . just the right amount of energy. to jump from one level to another. A . quantum. of energy is . just the right amount of energy. needed for an electron to jump levels.

Chapter 13 Electrons In Atoms

Chapter 13: Electrons in Atoms. Light is produced when electrons gain a small, specific amount of energy called a quantum. Once they gain a quantum of energy, they "leap" from their ground state, the lowest most stable state for an electron, to their excited state, a higher, less stable state for an electron.

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