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In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as we did

in Section 11.4 .

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Solution: (i) According to question, Centre of circle $(h, k) = (-2, 3)$ and Radius of circle $r = 4$ unit. Then, equation of circle. From formula $(x - h)^2 + (y - k)^2 = r^2$. $[x - (-2)]^2 + (y - 3)^2 = 4^2$. $\Rightarrow (x + 2)^2 + (y - 3)^2 = 16$. $\Rightarrow x^2 + 4x + 4 + y^2 - 6y + 9 = 16$.

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Solution: (i) According to question, Centre of circle $(h, k) = (-2, 3)$ and Radius of circle $r = 4$ unit. Then, equation of circle. From formula $(x - h)^2 + (y - k)^2 = r^2$. $[x - (-2)]^2 + (y - 3)^2 = 4^2$. $\Rightarrow (x + 2)^2 + (y - 3)^2 = 16$. $\Rightarrow x^2 + 4x + 4 + y^2 - 6y + 9 = 16$.

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