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Section 11.1 What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. 2. 3. 4. 5. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry. Stoichiometry is based on the law of conservation of mass,

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Name CHAPTER Section 11.2 continued Date Class STUDY GUIDE In your textbook, read about mole-to-mass and mass-to-mass conversions. Solving a mass-to-mass problem requires the four steps listed below.

CHAPTER Section 11.3 Limiting Reactants Date Class STUDY GUIDE In your textbook, read about why reactions stop and how to determine the limiting reactant. Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27

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In Section 11.3 , for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section ...

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11.1 Defining Stoichiometry MAIN Idea The amount of each reactant present at the start of a chemical reaction determines how much product can form. 11.2 Stoichiometric Calculations MAIN Idea The solution to every stoichiometric problem requires a balanced chemical equation. 11.3 Limiting Reactants MAIN Idea A chemical reaction

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