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Gas Laws STUDY GUIDE Due: February 12th Units of Measurement: For the following questions, use the following answer choices to indicate what each unit of measurement is used to measure. A. Pressure B. Volume C 1. K 2. kPa A 2. atm 3. L 3. mL 4. °C C. Temperature 'A 7. A 8.

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### Gases & Gas Laws Study Guide - Practice Test Questions ...

of moles of Ne (or any gas) present in a sample is directly proportional to the mass of the gas sample, the problem could also have been set up directly in terms of the masses. 21. Real gases most closely approach ideal gas behavior under conditions of relatively high temperatures (0°C or higher) and relatively low pressures (1 atm or lower). 22.

Study Guide for AP Chemistry – Chapter 5, Gas Laws . Students should be able

to... Explain the kinetic theory of matter as it applies to gases. Identify the properties of an ideal gas vs. a real gas . Know units of pressure in atm, torr, mm Hg, and kPa,

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190 Study Guide for An Introduction to Chemistry Section Goals and Introductions Section 13.1 Gases and Their Properties Goals To describe the particle nature of both real and ideal gases. To describe the properties of gases that can be used to explain their characteristics: volume, number of particles, temperature, and pressure.

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### Chemistry: Matter and Change Chapter 13 Gases Flashcards ...

### Chapter 14

### Study Guide for AP Chemistry - Chapter 5, Gas Laws

The familiar compound H<sub>2</sub>O provides the evidence that substances occur in three different physical classes called states of matter. At room temperature, H<sub>2</sub>O is a dense fluid called a liquid. When this liquid is chilled to 0°C, it changes to

a rigid solid. If the liquid is heated to 100°C, however, it abruptly expands to a tenuous fluid called vapor or gas.

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Chapter 14 concludes with a section on converting between volumes of gases reacting and produced in a chemical reaction. The main idea in this section is: 1) The ratio of volumes of gases in a reaction is the same as the ratio of moles, provided that the gas volumes are measured at the same temperature and pressure.

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### Chapter 14

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