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Quiz: Acids and Bases 1. Both acids and bases: A have a sour taste B contain hydrogen ions in solution C are corrosive 2. The pH of acids is: A greater than 7 B equal to 7 C less than 7 3. Bases turn litmus paper: A blue B purple C red 4. The name given to a soluble base is: A an acid B a ...

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an acid solution < 7; a base solution > 7; and; a neutral solution = 7. (c) The salts of strong acids and weak bases give acidic solution having pH less than 7. Example, NH₄Cl, Ammonium Chloride will have pH less than 7. The salts of weak acids and strong bases give basic solution having pH more than 7.

Acids, Bases & Oxides | CIE IGCSE Chemistry | MCQ & Answers

Tri acidic bases are those which have three 'OH' group per unit as Al(OH)₃ and Fe(OH)₃, their one mole requires three mole of a base (NaOH) for complete neutralization. 365 366 367 Answer

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Liquids that measure 0-7 are acids with 0 being the strongest. Bases will measure from 7-14 with 14 being the strongest. A pH of 7 means that the liquid is neutral. Distilled water has a pH of 7. The scale is a convention for measurement and some sources argue that acids can be negative figures and bases can be stronger than 14. [Acid and Base Worksheet - Answers](#) [Acids Bases and Salts multiple choice questions and ...](#)

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The chemistry of acids and bases and buffers is an important area. For example, the relative strengths of acids influences the formation of nitronium ions in the nitration of benzene and the understanding of pH and buffers is essential in biology. Very early in the history of chemistry, many substances were designated as acids, bases, and salts.

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Chapter 19 Acids Bases Salts Answer Key

2. Bases react with acids to produce salts and water; water to produce acids and salts; salts to produce acids and water; neither acids, salts, nor water

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Acid and Base Worksheet - Answers. 1) Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair: a) $HNO_3 + OH^- \rightarrow H_2O + NO_3^-$. HNO_3 and NO_3^- make one pair OH^- and H_2O make the other. b) CH_3NH_2

+ H_2O ($CH_3NH_3^+ + OH^-$)

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1 Acid = H_2O , base = NH_3 2 Acid = HCl, base = H_2O 3 Acid = $HCOOH$, base = KOH 4 Acid = HCl, base = CH_3COOH 5 Acid = HCl, base = NH_3 6 Acid = HCO_3^- , base = OH^-

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Acids and bases Acids, bases and alkalis are found in the laboratory and at home. Acids and bases can neutralise each other. A base that can dissolve in water is also called an alkali.

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2. Bases react with acids to produce salts and water; water to produce acids and salts; salts to produce acids and water; neither acids, salts, nor water

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an acid solution < 7 ; a base solution > 7 ; and; a neutral solution = 7. (c) The salts of strong acids and weak bases give acidic solution having pH less than 7. Example, NH_4Cl , Ammonium Chloride will have pH less than 7. The salts of weak acids and strong bases give basic solution having pH more than 7.

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Our general ideas regarding acid and alkali based on Arrhenius concepts in which an acid is a compound molecule that yield H^+ ion in water solution and the base is a compound molecule yields OH^- in water solution and utilization readily form solvent or water. Hence Arrhenius's theory very useful for explaining acid and bases in water solution but can't explain the strength in glacial acetic acid or ammonia solution.

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Acids and Bases - Definition, Examples, Properties, Uses ...

Acids Bases ; Arrhenius Concept: Acid is the substance when dissolved in water, increases the concentration of H^+ ions. The base is the substance when dissolved in water, increase the concentration of OH^- ions. Bronsted-Lowry Concept: Acids are the proton donor. Bases are the proton acceptor. Lewis Concept

Difference Between Acid and Base (with Comparison Chart ...)

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Question: Describe the properties of acids and bases My answer is acids are substances which form H^+ (aq) ions when

they are added to water. They are known as proton donors. Bases are also called alkalis, and are substances which form OH^- (aq) ions when added to water. They are known as proton acceptors. All acids have conjugate bases which are the particles that are left over when the acid ...

Our general ideas regarding acid and alkali based on Arrhenius concepts in which an acid is a compound molecule that yield H^+ ion in water solution and the base is a compound molecule yields OH^- in water solution and utilization readily form solvent or water. Hence Arrhenius's theory very useful for explaining acid and bases in water solution but can't explain the strength in glacial acetic acid or ammonia solution.

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Answers about Acids and Bases

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Question: Describe the properties of acids and bases My answer is acids are substances which form H^+ (aq) ions when they are added to water. They are known as proton donors. Bases are also called alkalis, and are substances which form OH^- (aq) ions when added to water. They are known as proton acceptors. All acids have conjugate bases which are the particles that are left over when the acid ...

Topic 12 - Acids, Bases and Buffers - A-Level Chemistry

4.3 Acids and Bases - A-Level Chemistry

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