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Acid - contains hydrogen and produces H + ions when dissolved in water. Base - contains hydroxide and produces OH - ions when dissolved in water. Binary acids (HF, HCl, H 2 S, etc.) Going across a row, rule is: the higher the electronegativity, stronger the acid. Ternary acids (HNO 3, H 2 SO 4, HClO 4 , etc.)

A strong acid is an acid which will dissociate completely to form an anion and a hydrogen ion. In contrast, a base is any type of molecule which can accept hydrogen ions from the solution. A ...

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Study Guide: Acids and Bases 1) List at least three characteristic properties of acids and three of bases. ACIDS pH less than 7 Have a sour taste Change the color of many indicators Are corrosive (react with metals) Neutralize bases Conduct an electric current BASES pH greater than 7

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In this summary sheet study guide, you will find the key concepts of acids and bases and acid-base reactions. It summarizes the factors such as electronegativity, polarizability, induction, and resonance-stabilization for determining the acidity of molecules and understating the p K a values and relationship to acid strength.

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Study Guide - Acids and Bases

Although the Arrhenius definitions of acid, base, and acid-base reactions provided in Sections 5.1, 5.4, and 5.6 are very important, especially to the beginning chemistry student, chemists have found it useful to extend these definitions to include new substances as acids and bases that would not be classified as such according to the Arrhenius definitions.

Chapter 5 Acids, Bases, and Acid-Base Reactions

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16.1 Acids and Bases: A Brief Review. •Acids taste sour and cause certain dyes to change color. •Bases taste bitter and feel soapy. •Arrhenius concept of acids and bases: •An acid is a substance that, when dissolved in water, increases the concentration of H+ions. •Example: HCl is an acid.

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Acids and Bases. An object's pH level can be tested using indicators. Objects with a low pH are acids, and those with a high pH are bases. Acids and bases react together to form water and salt.

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