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Section 3- Salts Neutralization- is a chemical reaction between an acid and a base that takes place in a water solution Salt- is a compound formed when the negative ions from an acid combine with the positive ions from a base.

Chemistry Notes - Acid, Bases and Salts - Chemistry ...

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• Acids and Bases react to form salt and water. Acid + Base →

Salt + H₂O • Neutralisation Reaction: Reaction of acid with a base is called as neutralization reaction. Example: HCl + NaOH → NaCl + H₂O • Strong Acid + Weak Base → Acidic salt + H₂O • Weak Acid + Strong Base → Basic salt + H₂O

Chapter - 2, Acids Bases and Salts Notes. • These are the substances which have sour in taste. • They turn blue litmus solution to red. • These substances are bitter in taste. • They turn red litmus solution to blue. • They give OH⁻ ions in aqueous solution. There are the substances that change their colour / smell in different types of substances.

Acids, Bases and Salts

In the crucial hour of Class 10 Board exams, Acids, Bases and Salts Class 10 Notes aims at increasing the self-confidence of the students by offering a simple way to study or revise the chapter. These notes provide the students with the summary of the chapter, important points to remember and detailed explanation of important concepts and derivations for better understanding.

Equations of Acids, Bases and Salts: Acid + Metal → Salt + Hydrogen gas H₂SO₄ + Zn → ZnSO₄ + H₂; Base + Metal → Salt + Hydrogen gas 2NaOH + Zn → Na₂ZnO₂ (Sodium zincate) + H₂; Base + Acid → Salt + Water NaOH (aq) + HCl (aq) → NaCl (aq) + H₂O (l) Acids give hydronium ions in water HCl + H₂O → H₃O⁺ + Cl⁻ Bases generate OH⁻ ions in water

Acid-Base Salt Note Taking Answers

Chapter 23 - Acids, Bases, and Salts

Note Taking Acids Bases And Salts Answers

Acid, Base And Salt Class 10 Science | Notes | Khullakitab

The salt formed by weak acid and strong base are called basic salt. E.g. Ca(HCO₃)₂. Methods of preparation of salts are as follows: Neutralization of acid and base NaOH + HCl → NaCl + H₂O. Reaction of acid on metallic oxides FeO + 2HCl → FeCl₂ + H₂O. Action of acids on metal Zn + H₂SO₄ → ZnSO₄ + H₂. The characteristics of salt are: Common salt (NaCl) is an essential constituent of our diet.

CBSE Class 10 Science Chapter 2 - Acids, Bases and Salts ...

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Add six drops of sulphuric (VI) acid /sodium sulphate (VI)/ammonium sulphate (VI) solution. Repeat with six drops of sodium sulphate (IV). (e) Reaction of cation with carbonate (IV)/CO₃²⁻ ions. All carbonate salts are insoluble except sodium/potassium carbonate (IV) and ammonium carbonate (IV).

Chapter-2, Acids Bases and Salts Notes

Revision Notes on Acids, Bases and Salts—Askiitians

a chemical reaction between an acid and a base taking place in a water solution. what is salt? ... Acids, Bases, and Salts packet. 40 terms. Unit 9: Acids, Bases, and pH. 27 terms. an acid is a polar substance that dissolves in water and. 27 terms. Science. OTHER SETS BY THIS CREATOR. 33 terms.

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neutralize, salts are formed. Strong acid and strong base combines to form neutral salt. $NaOH + HCl \rightarrow NaCl + H_2O$. Eq.1. Formation of Neutral Salt. Strong acid and weak base

Note Taking Acids Bases And Salts Answers

Note: The process of dissolving an acid or a base in water is a highly exothermic one. The acid must always be added slowly to water with constant stirring. Bases and Alkalis; A Base is a substance that gives OH^- ions when dissolved in water. Bases are usually metal hydroxides (MOH). According to Bronsted-Lowry theory, a base is a proton acceptor.

Study Notes On Chemistry: ACID, BASE AND SALTS

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Acid, Base And Salt Class 10 Science | Notes | Khullakitab

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Acids Bases and Salts Class 10 Notes Science Chapter 2 ...

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The salts of strong acids and strong bases are generally neutral. e.g. NaCl (from NaOH and HCl), K_2SO_4 (from KOH and H_2SO_4) The salts of strong acids and weak bases are acidic salts. e.g. $ZnSO_4$ (from H_2SO_4 and $Zn(OH)_2$) The salts of weak acids and strong bases are basic salts. e.g. Na_2CO_3 (from H_2CO_3 and NaOH)

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Hence, it may further take part in the reaction with the base as an acid. Such a salt is called an acid salt. For example, the salt $NaHSO_4$ produced in the reaction between NaOH and H_2SO_4 is an acid salt because it is capable of further reaction with the base NaOH to produce the normal salt Na_2SO_4 . $H_2SO_4 + NaOH \rightarrow NaHSO_4 + H_2O$...

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